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### **10 Chemical Quantities Guided Practice**

10.2 Mole-Mass and Mole-Volume Relationships 15. In your own words, describe how mass and moles are related. 16. Read sample problem 10.5 on page 298. Write out the equation that you would use to convert the number of moles of a substance to the mass of that substance. 17. Using dimensional analysis, answer Practice Problems 16-18 on p. 298-9.

### **Guided Notes Chapter 10: Chemical Quantities Name 10.1 The ...**

CHAPTER 10: Chemical Quantities BASICS: • The basic unit that

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is used to determine the amount of a chemical substance is called a mole • A mole(mol) of a substance is equivalent to  $6.02 \times 10^{23}$  particles of that substance • The mole was founded by a scientist named Avagadro, and he decided to use the

### **CHAPTER 10: Chemical Quantities**

Chapter 10 - Chemical Quantities. Jennie L. Borders. Section 10.1 - The Mole: A Measurement of Matter. ... Practice Problems. A compound formed when 9.03g Mg combines completely with 3.48g N. What is the percent composition of this compound? When a 14.2g sample of mercury (II) oxide is decomposed into its elements by heating, 13.2g of Hg is ...

### **Chapter 10 - Chemical Quantities**

92 Guided Reading and Study Workbook CHAPTER 10, Chemical Quantities(continued) 9. Complete the table about representative particles and moles. The Mass of a Mole of an

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Element (pages 293–294) 10. What is the atomic mass of an element? 11. Circle the letter of the phrase that completes this sentence correctly. The atomic masses of all elements

### **SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287–296)**

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Quantities91 SECTION 10.1 THE MOLE: A MEASUREMENT OF MATTER (pages 287–296) This section defines the mole and explains how the mole is used to measure matter. It also teaches you how to calculate the mass of a mole of any substance.  
Measuring Matter

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Prentice Hall Chemistry Chapter 10: Chemical Quantities Chapter Exam Instructions. Choose your answers to the questions and click 'Next' to see the next set of questions.

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Use the chemical formula to find the number of atoms in one molecule and multiply this number by Avogadro's number, the number of particles in one mole. atom 6.02 • 10<sup>23</sup> O 2 6.02 • 10<sup>23</sup> ion Na<sup>+</sup> 6.02 • 10<sup>23</sup> formula unit NaCl 6.02 • 10<sup>23</sup> 6.02 •

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1023 representative particles of a substance molecule formula unit atom

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