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Title: Colligative
Properties of Solutions
1 Colligative Properties
of Solutions. Boiling Pt
Elevation and Freezing
Pt Depression; 2
Colligative Properties.
depend only on the
number of solute
particles in a solution ;

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does not depend on the identity of particles; 3 Boiling Point Elevation. a nonvolatile solute lowers the vapor pressure

PPT - Colligative Properties of Solutions PowerPoint ...

Colligative Properties - Solute particles take up space in a solution. ... (density 0.88 g cm⁻³) in the osmometer

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corrected for capillary
rise is 11.6 cm at ... |
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Properties - $P_A = .848$ (
760 torr) $P_A = 644$ torr
...

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Colligative Properties
Colligative Property: A
property that depends
only upon the number

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of solute particles
(concentration), and
not upon their identity.
Three Important
Colligative Properties
of Solutions. Vapor-
pressure lowering
Boiling-point elevation
Freezing-point
depression

Chapter 16 **Colligative** **Properties of** **Solutions**

|author**STREAM**

1. Solutions Colligative

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Properties • Changes in colligative properties depend only on the number of solute particles present, not on the identity of the solute particles. • Among colligative properties are Vapor pressure lowering Boiling point elevation Melting point depression Osmotic Pressure. 2.

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Colligative properties are those that depend on the concentration of particles in a solution, not upon the identity of those particles.

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operating

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clearly built to give step-

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to go ahead in

operating certain

equipments.

PPT ON COLLIGATIVE PROPERTIES OF SOLUTIONS PDF

- By definition a

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colligative property is a solution property (a property of mixtures) for which it is the amount of solute dissolved in the solvent matters but the kind of solute does not matter.

- Coming to grips with this concept should immediately remind you of kinetic molecular theory of gases—in that case we

**Colligative
Properties- Page 1**

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COLLIGATIVE

PROPERTIES Elevation
of Boiling Point

Depression of Freezing
Point Lowering of

Vapor Pressure

Osmotic Pressure

MOLE FRACTION &

MOLALITY MOLE

FRACTION OF

Component $i = X_i = n_i$
/ n_{total} (c.f Gases;

Chapter 5, p.217)

MOLALITY = Moles of
Solute / kg Solvent

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MOLALITY Useful when
Temperature Changes
are considered, as
Volumes of solutions
change with changing
temperature, whereas
...

COLLIGATIVE PROPERTIES

Among colligative
properties are Vapor
pressure lowering
Boiling point elevation
Melting point
depression Osmotic
pressure © 2009,
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Vapor Pressure Ppt File

Because of solute-solvent intermolecular attraction, higher concentrations of nonvolatile solutes make it harder for solvent to escape to the vapor phase. ©

2009, Prentice-Hall, Inc. Raoult's Law $P_A = X_A P_A^*$ where X_A is the mole fraction of compound A, and P_A^* is the normal vapor pressure of A at that

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temperature.

Chapter 13
Properties of
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College

Properties of solutions that depend on the number of molecules present and not on the kind of molecules are called colligative properties. These properties include boiling point elevation, freezing point depression, and

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osmotic pressure.

Colligative Properties - University Of Cincinnati

Colligative Properties
Of Solutions. 1.

Colligative Properties
of Solutions
. 2.

The presence of
solutes affects the
properties of
solutions.
Some
of these properties are
not dependant on what
dissolved substance

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Properties Of

but only on how
much.
Colligative
properties are those
that depend on the
concentration of solute
particles but not on
their identity.<br
Properties?

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Colligative Properties
of Ionic Solutions
Colligative properties

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of solutions depends upon the total concentration of particles. Each

equation describing colligative properties must be modified to account for this with ionic solutions since each ionic compound gives more than one mole of ions for every mole of compound.

PowerPoint Presentation

Osmotic pressure is the

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most important of the colligative properties since it is related to the physiological compatibility of parenteral, ophthalmic and nasal solutions.

A Lecture on Colligative Properties in an Undergraduate ...

Colligative properties of solutions are properties that depend upon the concentration of solute molecules or

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ions, but not upon the identity of the solute.

Colligative properties include vapor pressure lowering, boiling point elevation, freezing point depression, and osmotic pressure.

Lowering the Vapor Pressure:

Colligative Properties - Chemistry & Biochemistry

Colligative Properties
Properties of solutions

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(compared to the pure solvent) Depend upon the number of solute molecules or ions (concentration), but NOT upon the identity of the solute. Include freezing point depression, boiling point elevation, vapor pressure, and osmotic pressure. You should be able to: 1. Define the CP 2.

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Colligative Properties

A. Definition Colligative

Property • property

that depends on the

number (amount) of

solute particles in the

solution & not in the

nature of the solute

particles Colligative

Properties Vapor

Pressure Lowering

Boiling Point Elevation

Freezing Point

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Depression Osmotic
Pressure Vapor
Pressure Lowering
Vapor Pressure-
pressure exerted by
the vapor/gas above
the liquid VP of sol'n <
VP of solvent $P = P_1^0 -$
 $P_1 = X_2 P_1^0$ where: P_1^0
= V.P. of solvent $X_2 =$
mole frxn of ...

colligative.ppt - Solutions I II III III Colligative ...

Colligative Properties A
dilute solution is one in

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which the amount of the solute is very small in comparison to the amount of the solvent. The dilute solutions show more or less ideal behavior as the heat and volume changes, accompanying the mixing of solute and solvent, are negligible for all practical purposes.

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