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S-14 1. Solubility of Ethanol in Water Explain why ethanol ($\text{CH}_3\text{CH}_2\text{OH}$) is more soluble in water than is ethane (CH_3CH_3). Answer Ethanol is polar; ethane is not. The ethanol $-\text{OH}$ group can hydrogen-bond with water. 2. Calculation of pH from Hydrogen Ion Concentration What is the pH of

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a solution that has an H concentration of (a) 1.75×10^{-5} mol/L; (b) 6.50×10^{-10} mol/L; (c) 1.0×10^{-4} mol/L;

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