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The first law of thermodynamics states: In a process without transfer of matter, the change in

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internal energy,  $\Delta U$ , of a thermodynamic system is equal to the energy gained as heat,  $Q$ , less the thermodynamic work,  $W$ , done by the system on its surroundings.  $\Delta U = Q - W$ .  $\{\displaystyle \Delta U=Q-W\}$  .

### **Thermodynamics - Wikipedia**

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